

Date Planned : __ / __ / __	Daily Tutorial Sheet-4	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	Level-1	Exact Duration : _____

46. In halogen's group which element has highest electron affinity ?
(A) F (B) Cl (C) Br (D) I
47. Which one is the anhydride of HClO_4 ? 
(A) ClO_2 (B) Cl_2O_7 (C) Cl_2O (D) Cl_2O_6
48. In which case, the order of acidic strength is not correct ?
(A) $\text{HI} > \text{HBr} > \text{HCl}$ (B) $\text{HIO}_4 > \text{HBrO}_4 > \text{HClO}_4$
(C) $\text{HClO}_4 > \text{HClO}_3 > \text{HClO}_2$ (D) $\text{HF} > \text{H}_2\text{O} > \text{NH}_3$
49. Correct order of electron affinities of halogens is :
(A) $\text{F} > \text{Cl} > \text{Br} > \text{I}$ (B) $\text{I} > \text{Br} > \text{Cl} > \text{F}$ (C) $\text{Cl} > \text{F} > \text{I} > \text{Br}$ (D) $\text{Cl} > \text{F} > \text{Br} > \text{I}$
50. The following species will not exhibit disproportionation reaction : 
(A) ClO^- (B) ClO_2^- (C) ClO_3^- (D) ClO_4^-
51. The correct order of heat of formation of halogen acids is :
(A) $\text{HI} > \text{HBr} > \text{HCl} > \text{HF}$ (B) $\text{HF} > \text{HCl} > \text{HBr} > \text{HI}$
(C) $\text{HCl} > \text{HF} > \text{HBr} > \text{HI}$ (D) $\text{HCl} > \text{HBr} > \text{HF} > \text{HI}$
52. Which of the following pairs has bleaching property ?
(A) O_3 and NO (B) O_3 and H_2S (C) SO_2 and Cl_2 (D) H_2 and NO_2
53. Which of the following is mixed anhydride ? 
(A) Cl_2O_7 (B) ClO_2 (C) Cl_2O_3 (D) Cl_2O
54. Most acidic oxide among the following is :
(A) ClO_2 (B) Cl_2O (C) Cl_2O_3 (D) Cl_2O_7
55. Iodine readily dissolves in potassium iodine solution giving : 
(A) I^- (B) KI^- (C) KI_2^- (D) KI_3
56. Least stable oxide of chlorine is :
(A) Cl_2O (B) ClO_2 (C) Cl_2O_7 (D) ClO_3
57. The deficiency of iodine in diet causes :
(A) Rickets (B) Scurby (C) Beri-beri (D) Goiter
58. The weakest acid is :
(A) HI (B) HBr (C) HCl (D) HF
59. Which one will liberate Br_2 from KBr ?
(A) I_2 (B) SO_2 (C) HI (D) Cl_2
60. The bond dissociation energy of Cl_2 , Br_2 and I_2 follow the order : 
(A) $\text{Cl}_2 > \text{I}_2 > \text{Br}_2$ (B) $\text{I}_2 > \text{Br}_2 > \text{Cl}_2$ (C) $\text{I}_2 = \text{Cl}_2 = \text{Br}_2$ (D) $\text{Cl}_2 > \text{Br}_2 > \text{I}_2$